

Assignment \_\_\_\_\_ Name \_\_\_\_\_ Hour \_\_\_\_\_

### Covalent Bonding Worksheet 2

1. Read Chapter 9, Sections 2 and 3 pages 248-258.
2. Draw an orbital notation for a **hydrogen fluoride molecule**.
3. Draw an electron dot notation for a **hydrogen fluoride molecule**.
4. Draw the structural formula for a **hydrogen fluoride molecule**.
5. The molecular formula for a **hydrogen fluoride molecule** is \_\_\_\_\_.
6. Draw an orbital notation for a **hydrogen sulfide molecule**.
7. Draw an electron dot notation for a **hydrogen sulfide molecule**.
8. Draw the structural formula for a **hydrogen sulfide molecule**.
9. The molecular formula for a **hydrogen sulfide molecule** is \_\_\_\_\_.
10. How many electrons do two atoms in a double covalent bond share? How many electrons do two atoms in a triple covalent bond share?

11. The following covalent molecules have only single covalent bonds. Draw an electron dot structure for each.

